



FORMULAE & EQUATIONS (A)

- 1** Give the formula of the following substances.
- a** potassium carbonate (1)
 - b** magnesium hydroxide (1)
 - c** oxygen (1)
 - d** aluminium bromide (1)
 - e** sodium (1)
 - f** argon (1)
 - g** iron(III) sulfate (1)
 - h** phosphorus (1)
 - i** copper(I) oxide (1)
 - j** hydrogen sulfide (1)
- 2** In what molar ratio do the following substances react?
- a** hydrochloric acid with barium hydroxide (1)
 - b** sulfuric acid with ammonia (1)
 - c** nitric acid with sodium hydrogencarbonate (1)
- 3** Write an ionic equation, including state symbols, for each of the following reactions.
- a** reaction of aqueous potassium carbonate with nitric acid
..... (2)
 - b** precipitation of lead(II) bromide when aqueous lead(II) nitrate is mixed with aqueous sodium bromide
..... (2)
 - c** reaction of aqueous ammonia with sulfuric acid
..... (2)
 - d** reaction of hydrochloric acid with aqueous potassium hydroxide
..... (2)