



- 1 a In what order are the elements arranged in the Periodic Table? **atomic number order**
- b How many electrons are in the outer shell of atoms of the following elements?  
aluminium **3**    fluorine **7**    silicon **4**
- c Give the group and period number of the element with electron structure 2,8,5.  
group **5**    period **3**
- d Which group are the following elements in? The electron structure of these elements is given.  
2,8,8,1 **1**    2,6 **6**    2,8,18,5 **5**
- 2 a Explain why elements that are in the same group in the Periodic Table have similar properties.  
**same number of electrons in the outer shell**
- b Explain why the elements in Group 0 are unreactive.  
**they have stable electron structures and so do not need to gain/lose/share electrons**
- c Explain why the elements in Group 1 are very reactive.  
**they have one electron in their outer shell and so it is easy to lose one electron to get a stable electron structure**