



1 Give the formula of the following ions.

bromide sulfate silver(I)

2 Give the formula of the following ionic compounds.

potassium sulfide sodium carbonate

calcium hydroxide aluminium nitrate

3 Potassium fluoride is an ionic compound containing K^+ and F^- ions.

a Give the electron structure of the K^+ ions.

b Give the electron structure of the F^- ions.

c Potassium fluoride melts at 858°C . Explain why potassium fluoride has a high melting point.

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d Explain why potassium fluoride conducts electricity when molten.

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e Explain why potassium fluoride does not conduct electricity as solids.

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