



QUICK CHECK

THERMODYNAMICS (A)

Write a chemical equation that represents each of the following enthalpy changes.

a enthalpy of formation of H₂O(l)



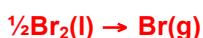
b enthalpy of combustion of C₂H₆(g)



c enthalpy of vaporisation of H₂O(l)



d enthalpy of atomisation of Br₂(l)



e lattice enthalpy of formation of MgCl₂(s)



f lattice enthalpy of dissociation of CaO(s)



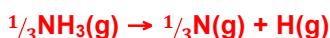
g enthalpy of solution of MgCl₂(s)



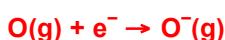
h enthalpy of hydration of Mg²⁺ ions



i bond dissociation enthalpy for N-H in NH₃



j 1st electron affinity of oxygen



k 2nd electron affinity of oxygen



l 2nd ionisation enthalpy of magnesium

