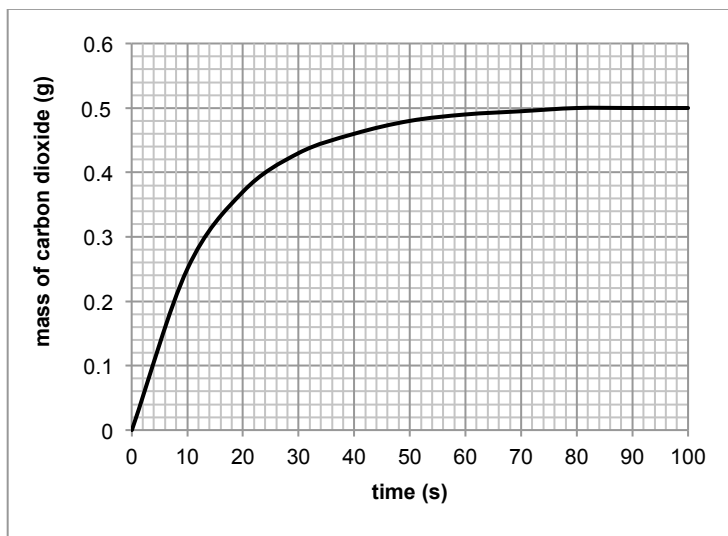
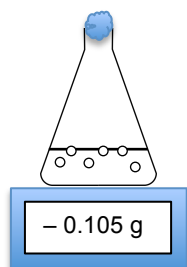




A student carried out an experiment where he recorded the mass of carbon dioxide gas formed as calcium carbonate reacts with hydrochloric acid.

The reaction was done on a set of scales. The student tared (zeroed) the scales at the start of the reaction so he could measure the mass of gas lost. He plotted the mass of gas lost against time on a graph.



- 1 Why is cotton wool placed in the neck of the conical flask? .....
- 2 Explain why the rate slows down and then stops. ....
- 3 Calculate the mean rate of reaction in the first 10 seconds in g/s .....
- 4 Draw a tangent to the graph to find the rate at 10 seconds in g/s .....
- 5 Explain why this reaction would be faster if more a more concentrated solution of acid was used. ....