



1 a What is a Bronsted-Lowry acid?

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b Identify the Bronsted-Lowry acid in this reaction. $\text{H}_2\text{O} + \text{CH}_3\text{NH}_2 \rightarrow \text{OH}^- + \text{CH}_3\text{NH}_3^+$

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2 a Define pH.

b Calculate the pH of a solution of nitric acid with concentration $0.200 \text{ mol dm}^{-3}$.

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c Calculate the concentration of a solution of sulfuric acid with pH 1.30.

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d Calculate the pH of the solution formed when 200 cm^3 of water is added to 50 cm^3 of $0.800 \text{ mol dm}^{-3}$ hydrochloric acid.

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