



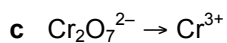
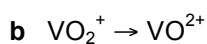
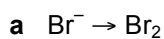
1 Identify the oxidising agent and reducing agent in this reaction: $F_2 + 2KBr \rightarrow 2KF + Br_2$

oxidising agent = reducing agent =

2 State the oxidation state of the sulfur in the following species.

species	H ₂ S	S ₈	SO ₃ ²⁻	H ₂ SO ₄	SO ₂	SF ₆	NaHSO ₃
oxidation state of S							

3 Write a half equation for each of the following conversions.



4 Combine these pairs of half equations to make a redox equation.

