

	1	2	3	4	5	6
1	relative atomic mass	<b>Avogadro constant</b>	empirical formula	?	<b>relative molecular mass</b>	<i>ionic equation</i>
2	<i>Percentage atom economy</i>	% yield	?	<i>cm<sup>3</sup> to m<sup>3</sup></i>	High value holds ethical, economic and environmental advantages for society	<i>...simplest whole number ratio of the atoms of each different element in a given compound...</i>
3	<i>pV = nRT</i>	<b>One electron has a mass of 9.10938291 x 10<sup>-31</sup> kg.</b>	$\frac{\text{Molecular mass of desired product}}{\text{Sum of molecular masses of all reactants}} \times 100$		?	
4	?	<b>Calculate the mass of 1 mole of electrons.</b>	<i>A<sub>r</sub></i>	<i>water of crystallisation</i>	<b>volumetric solution</b>	<i>6.022 x 10<sup>23</sup></i>
5	<i>mol dm<sup>-3</sup></i>	<b>...actual number of atoms of each element in a compound...</b>	?	<b>actual mass/theoretical mass x100%</b>	<i>the ratio of the average mass of one atom of an element to one twelfth of the mass of an atom of carbon-12</i>	<i>M<sub>r</sub></i>
6	empirical formula	<b>CH<sub>2</sub> and C<sub>5</sub>H<sub>10</sub></b>	$n = \frac{pV}{RT}$	<i>ideal gas equation</i>	<i>molecular formula</i>	?

DP Amount of Substance