	1	2	3	4	5	6
1	relative atomic mass	Avogadro constant	empirical formula	Kelvin	relative molecular mass	*ionic equation*
2	*Percentage atom economy*	% yield	titration	cm^3 to m^3	High value holds ethical, economic and environmental advantages for society	…simplest whole number ratio of the atoms of each different element in a given compound…
3	$pV = nRT$	One electron has a mass of $9.10938291 \times 10^{-31}$ kg.	$\frac{\text{Molecular mass of desired product}}{\text{Sum of molecular masses of all reactants}} \times 10$		*mole*	
4	*mass / relative mass*	Calculate the mass of 1 mole of electrons.	A_r	water of crystallisation	volumetric solution	6.022×10^{23}
5	$mol\ dm^{-3}$	…actual number of atoms of each element in a compound…	standard form	actual mass/ theoretical mass x100%	the ratio of the average mass of one atom of an element to one twelfth of the mass of an atom of carbon-12	M_r
6	empirical formula	CH_2 and C_5H_{10}	$n = \frac{pV}{RT}$	ideal gas equation	molecular formula	m^3 to dm^3

DP Amount of Substance