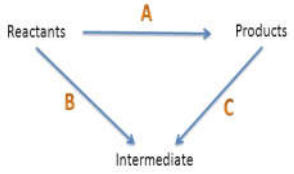
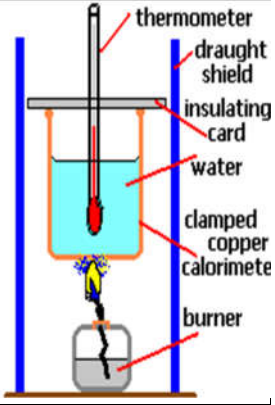
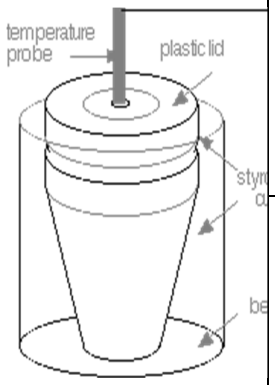
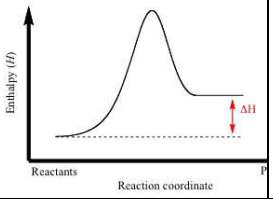
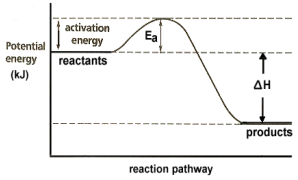
	1	2	3	4	5	6
1	kJ mol^{-1}	calorimetry	Hess's Law	Q = mcΔT	Mean bond enthalpy	$C_2H_5OH_{(l)} + 3O_{2\,(g)} \rightarrow 2CO_{2\,(g)} + 3H_2O_{(l)}$
2	*Total enthalpy change for a reaction is independent of route taken*	ΔH	ΔH_f		ΔH is negative	
3	**standard conditions**		$\Sigma H_{total} = \Sigma H_{reactants} - \Sigma H_{products}$		*Heat energy transferred in a reaction at constant pressure*	
4	*bond enthalpy*		ΔH = breaking bonds – forming bonds	$2C_{(s)} + 3H_{2\,(g)} + O_{2\,(g)} \rightarrow C_2H_5OH_{(l)}$	*100 kPa and 298 K*	*Joules*
5	ΔH_r		q/n	Pa	*The enthalpy change when 1 mole of a substance is completely burned in oxygen, under standard conditions, with all reactants and products in their standard states*	
6	Specific heat capacity	***endothermic***		***Bomb calorimeter***	ΔH_c	*exothermic*

DP Energetics