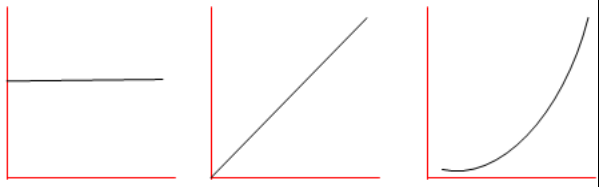
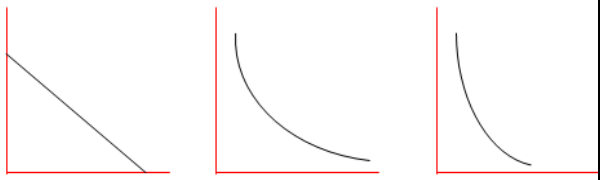
	1	2	3	4	5	6
1	?	Arrhenius constant	initial rate method	$\ln k = -E_a/RT + \ln A$	second-order reaction	*rate equation*
2	*activation energy*	the rate of reaction is proportional to the concentration of reactant A	straight line graph of slope = $-E_a/R$	?	effect of catalyst on value of rate constant	reaction mechanism
3	zero-order reaction	*continuous monitoring method*			?	*concentration-time graph*
4	*gradient*	reaction order	?	$X + Y \rightleftharpoons Z$, rate = k[Y]	*initial rate of a reaction*	*rate determining step*
5	*mathematical relationship between rate constant, temperature and activation energy*	$k = Ae^{-E_a/RT}$	*first-order reaction*	plot 1/t against concentration	*Arrhenius equation*	?
6	$[CH_3CO_2H]$	?	*effect of temperature on rate constant?*	*rate-concentration graphs*		

Rate Equations