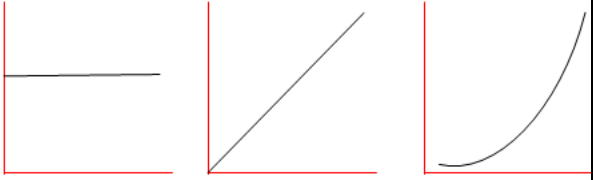
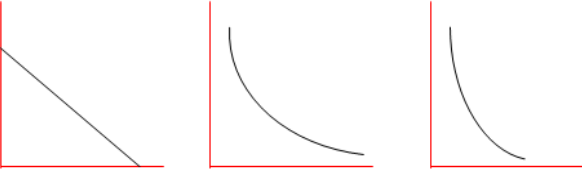


	1	2	3	4	5	6
1	rate equation	Arrhenius constant	initial rate method	$\ln k = -E_a/RT + \ln A$	second-order reaction	rate constant
2	<i>activation energy</i>	the rate of reaction is proportional to the concentration of reactant A	straight line graph of slope = $-E_a/R$	$[A]^2$	effect of catalyst on value of rate constant	reaction mechanism
3	zero-order reaction	<i>continuous monitoring method</i>			$rate = k[A][B]^2$	<i>concentration-time graph</i>
4	<i>gradient</i>	reaction order	<i>units?</i>	$X + Y = Z,$ $rate = k[Y]$	<i>initial rate of a reaction</i>	<i>rate determining step</i>
5	mathematical relationship between rate constant, temperature and activation energy	$k = Ae^{-E_a/RT}$	first-order reaction	plot 1/t against concentration	<i>Arrhenius equation</i>	<i>rate of reaction</i>
6	$[CH_3CO_2H]$	<i>what if a species involved in a reaction doesn't appear in the rate equation?</i>	<i>effect of temperature on rate constant?</i>	<i>rate-concentration graphs</i>		

Rate Equations