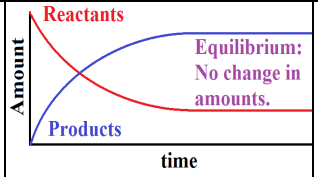
	1	2	3	4	5	6
1	Increase pressure on $2SO_2 + O_2 \rightleftharpoons 2SO_3$ system	$C_2H_{4(g)} + H_2O_{(g)} \rightleftharpoons C_2H_5OH_{(g)}$	?	**Exothermic reaction**	**Reversible reaction**	$K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$
2	*'the position of equilibrium will move in such a way as to counteract the change'*	*Dynamic equilibrium*	Haber Process	***Speeds up forward and backward reaction equally***	?	**Compromise temperature**
3	$N_2 + 3H_2 \rightleftharpoons 2NH_3$ $\Delta H = -92kJ\ mol^{-1}$	↑ temperature on $2SO_2 + O_2 \rightleftharpoons 2SO_3$ system ($\Delta H = -197$ kJ mol^{-1})	*Heterogeneous equilibrium*	***Decrease concentration of NO_2 in $N_2O_{4(g)} \rightleftharpoons 2NO_{2(g)}$ system***	*Position of equilibrium*	?
4	?	**Forward and backward reactions proceed at equal rates**	Rate of reaction versus yield	Decrease pressure on $H_{2(g)} + I_{2(g)} \rightleftharpoons 2HI_{(g)}$ system	***Equilibrium constant***	*Le Chatelier's principle*
5	*400 – 450 °C 200 atm Fe catalyst*		Backwards reaction	?	*Compromise temperature*	*Endothermic reaction*
6	mol dm^{-3}	?	*Kc and Catalysts*	*Homogeneous reaction*	$2H_{2(g)} + CO_{(g)} \rightleftharpoons CH_3OH_{(g)}$	***Closed system***

DP Equilibrium