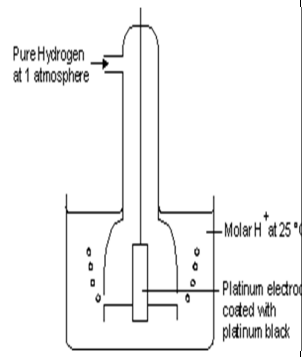
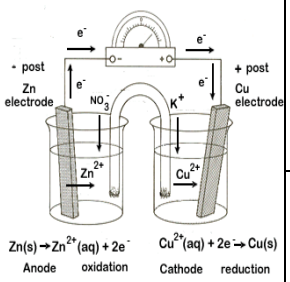
	1	2	3	4	5	6
1	*half-equations for electrode reactions*	**cell EMF**	$E^{\ominus}$	$Zn\|Zn^{2+}\|\|Cu^{2+}\|Cu$	**rechargeable cell**	**298 K, 100 kPa and 1.00 mol dm^{-3} solution of ions**
2	*fuel cell*	electrode potentials	**alkaline hydrogen–oxygen fuel cell**	***how can electrode reactions be used to generate an electrical current?***	Positive electrode: $Li^+ + CoO_2 + e^- \rightarrow Li^+[CoO_2]^-$ Negative electrode: $Li \rightarrow Li^+ + e^-$	
3	***predicting the direction of simple redox reactions***	*e.g. mobile phones, laptops or tablets*	reference cell	standard electrode potential	*definition of reducing agent*	***benefits and risks to society associated with using different types of cells***
4	*strongest oxidising agent*	electrochemical series	feasibility of a reaction	**the equilibrium with the more negative E° value will move to the left**	***electrochemical cells as a commercial source of electrical energy***	
5	*Pt electrodes*	non-rechargeable cell		**salt bridge**	weakest reducing agent	
6	***importance of the conditions when measuring the electrode potential, E***	***lithium cell***		Li | Li+ || Li+, COO2 | LiCoO2 | Pt	*standard hydrogen electrode (she)*	most positive standard electrode potential

Electrode Potentials