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What's the Question?

What is "What's the question?"?

This game is a Mole Day variation of the US TV programme Jeopardy.

The aim of the game is to get the most points. Points are gained by asking the question for which the answer is given.

There are five categories (IUPAC rules, gases, solutions, liquids and solids). Each category contains 5 questions. Harder questions are worth more points (from 1 – 5).

Starting the game

Give the clue "An Italian lawyer whose law states that equal volumes of gases at the same temperature and pressure contain an equal number of molecules"

The first student to have their hand up and give the correct answer in the form of a question, e.g. "Who is Amedeo Avogadro?" chooses the first question.

To play

Split the class into teams for point scoring purposes, although they will answer questions as individuals and pool answers at the end.

During the game, players can choose any unselected clue, e.g. "solutions for 5".

Any student can put their hand up to answer any clue.

The correct answer must be given in the form of a question.

The student who answers this first question correctly chooses the next question by stating the category and point value of question, e.g. "IUPAC rules for 5" (the hardest question on IUPAC rules), which is again open to anyone.

When a student answers a question correctly, they are given the number of points corresponding to the question they give.

Keep track of points by writing the name of the student who answered the question correctly on the point sheet.

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IUPAC DEFINITIONS	GASES	SOLUTIONS	LIQUIDS	SOLIDS
<p>5</p> <p>The ratio of the average mass of the atom to the unified atomic mass unit.</p> <p><i>Ans: What is relative atomic mass?</i></p>	<p>5</p> <p>Amount of one gas in a mixture divided by the total amount of gas in the mixture.</p> <p><i>Ans: What is mole fraction?</i></p>	<p>5</p> <p>A solution of a known concentration.</p> <p><i>Ans: What is a standard solution?</i></p>	<p>5</p> <p>A substance that conducts electric current when molten or dissolved in water.</p> <p><i>Ans: What is an electrolyte?</i></p>	<p>5</p> <p>Mass in grams divided by formula mass</p> <p><i>Ans: What is amount of substance?</i></p>
<p>4</p> <p>Name given to the mass of 1 mole of substance</p> <p><i>Ans: What is molar mass?</i></p>	<p>4</p> <p>The volume occupied by 2 mol CO₂ at room temperature and 1 atm pressure</p> <p><i>Ans: What is 48 dm³</i></p>	<p>4</p> <p>Amount concentration, given by amount of substance divided by the volume of the solution.</p> <p><i>Ans: What is molarity?</i></p>	<p>4</p> <p>The temperature at which a liquid changes state to become a gas.</p> <p><i>Ans: What is boiling point?</i></p>	<p>4</p> <p>A radioactive metal discovered by Marie and Pierre Curie and used by Ernest Rutherford to propose a new model of atomic structure.</p> <p><i>Ans: What is radium?</i></p>
<p>3</p> <p>6.02×10^{23}</p> <p><i>Ans: What is Avogadro's number?</i></p>	<p>3</p> <p>Equal volumes of gases at the same temperature and pressure contain an equal number of molecules</p> <p><i>Ans: What is Avogadro's law?</i></p>	<p>3</p> <p>The unit for concentration.</p> <p><i>Ans: What is mol dm⁻³?</i></p>	<p>3</p> <p>Ingesting this metal with molar mass 201 g mol⁻¹ will make you "mad as a hatter"</p> <p><i>Ans: What is mercury?</i></p>	<p>3</p> <p>The second most abundant element in the Earth's crust with a molar mass of 28 g mol⁻¹, this element is needed to make solar cells</p> <p><i>Ans: What is silicon?</i></p>
<p>2</p> <p>SI base unit for "amount of substance"</p> <p><i>Ans: What is the mole?</i></p>	<p>2</p> <p>273K, 1 atm pressure</p> <p><i>Ans: What is standard temperature and pressure?</i></p>	<p>2</p> <p>1 dm³ or 1000 cm³</p> <p><i>Ans: What is 1 litre?</i></p>	<p>2</p> <p>The molar mass of this common liquid is 18 g mol⁻¹</p> <p><i>Ans: What is water?</i></p>	<p>2</p> <p>One mole of substance contains the same number of elementary entities as there are atoms in 0.012 kg of this solid element.</p> <p><i>Ans: What is carbon 12?</i></p>
<p>1</p> <p>Smallest particle still characterising a chemical element</p> <p><i>Ans: What is an atom?</i></p>	<p>1</p> <p>Approximate volume of one mole of gas at 298 K and 1 atmosphere pressure.</p> <p><i>Ans: What is 24 dm³?</i></p>	<p>1</p> <p>The state symbol that shows that the physical state of a chemical in an equation is an aqueous solution?</p> <p><i>Ans: What is (aq)?</i></p>	<p>1</p> <p>The mass of a liquid divided by its volume</p> <p><i>Ans: What is density?</i></p>	<p>1</p> <p>The SI unit of mass.</p> <p><i>Ans: What is the kilogram?</i></p>

LINKING LANGUAGE



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Score board

IUPAC RULES	GASES	SOLUTIONS	LIQUIDS	SOLIDS
5	5	5	5	5
4	4	4	4	4
3	3	3	3	3
2	2	2	2	2
1	1	1	1	1

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